Ciências
ULisboa

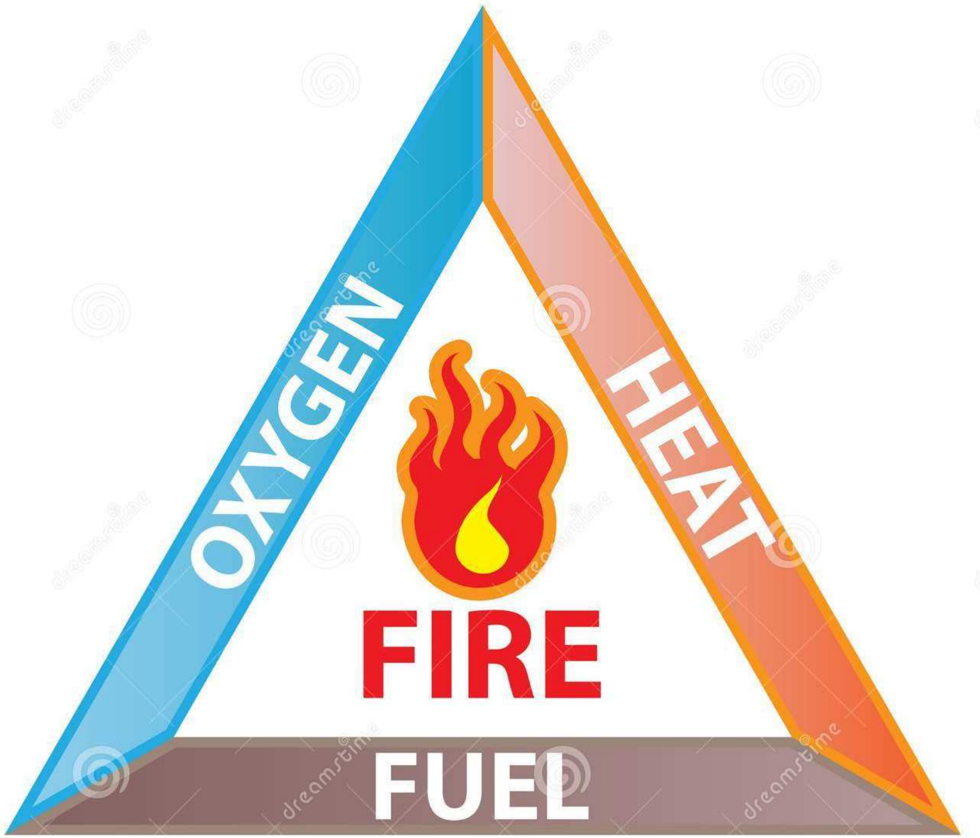
Faculdade
de Ciências
da Universidade
de Lisboa

DISCIPLINA MIEEA 2018

Technologies of combustion

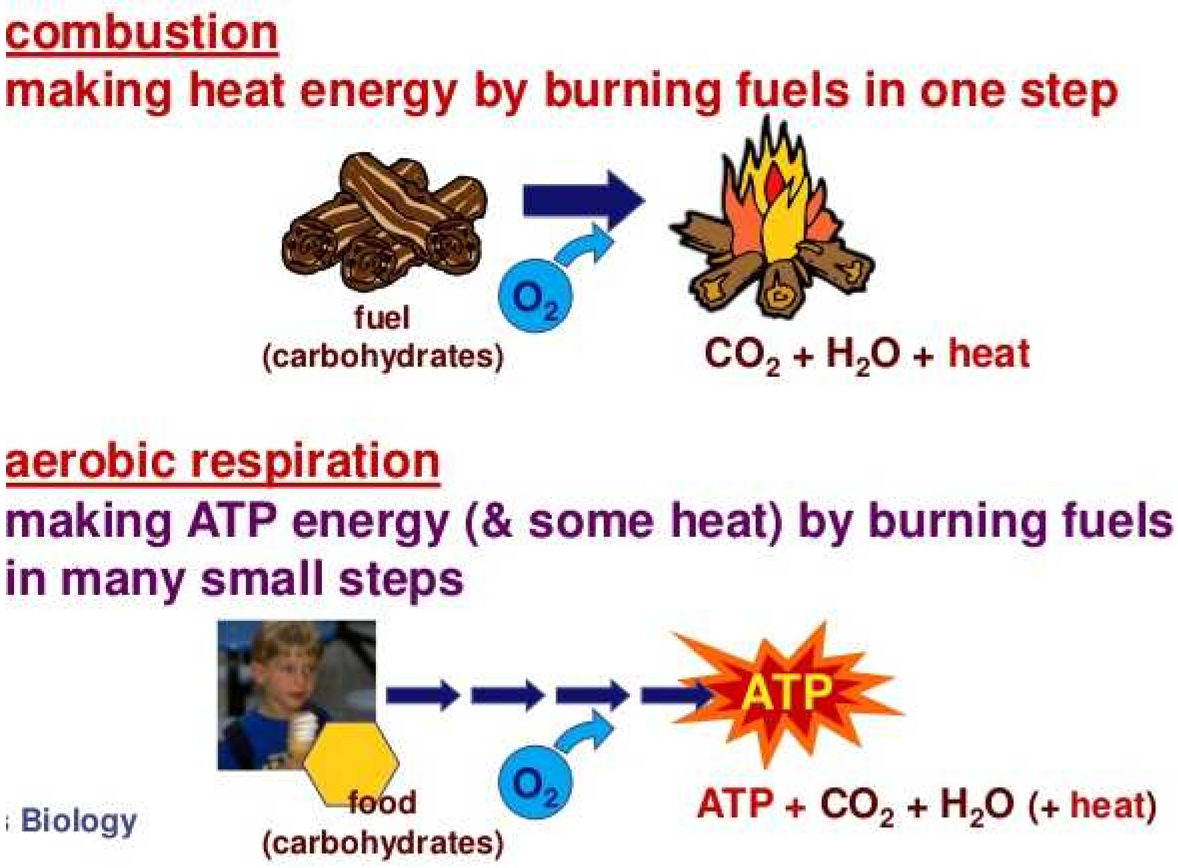
Combustion definition

Combustion is essentially burning, fuels react with oxygen to release energy

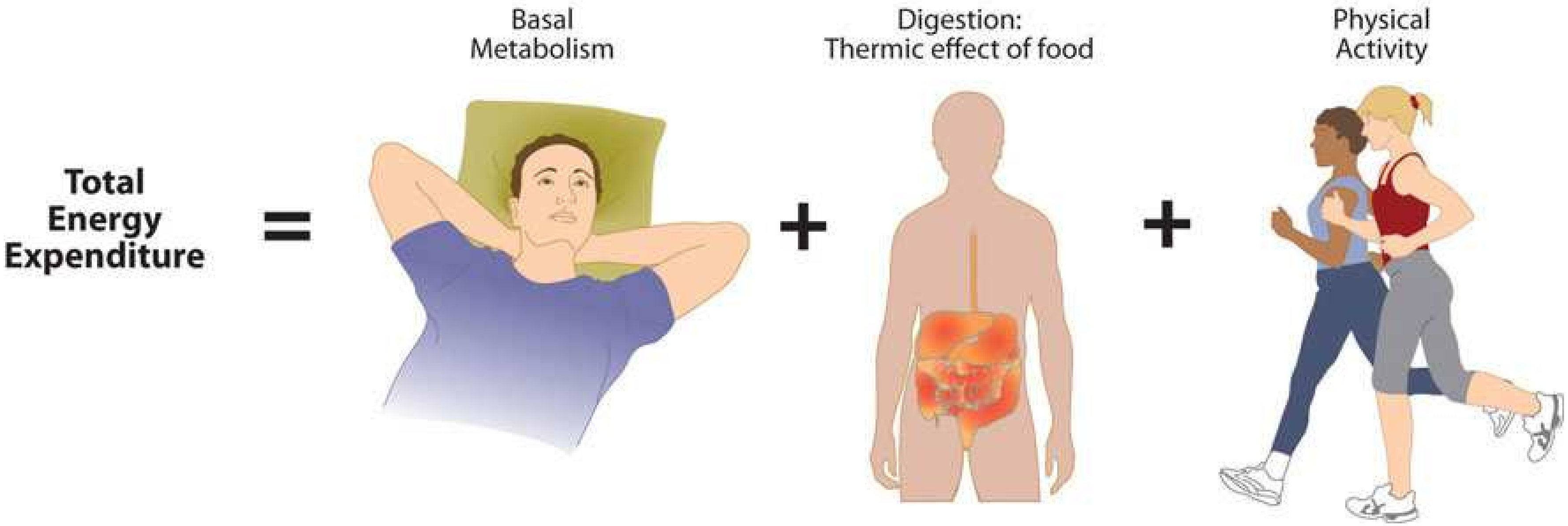


Combustion in biology

Respiration is the chemical reaction in which energy is released from a reaction between Oxygen (O_2) and Glucose ($C_6H_{12}O_6$) Respiration releases energy for cells from glucose



Combustion in biology



Combustion in biology

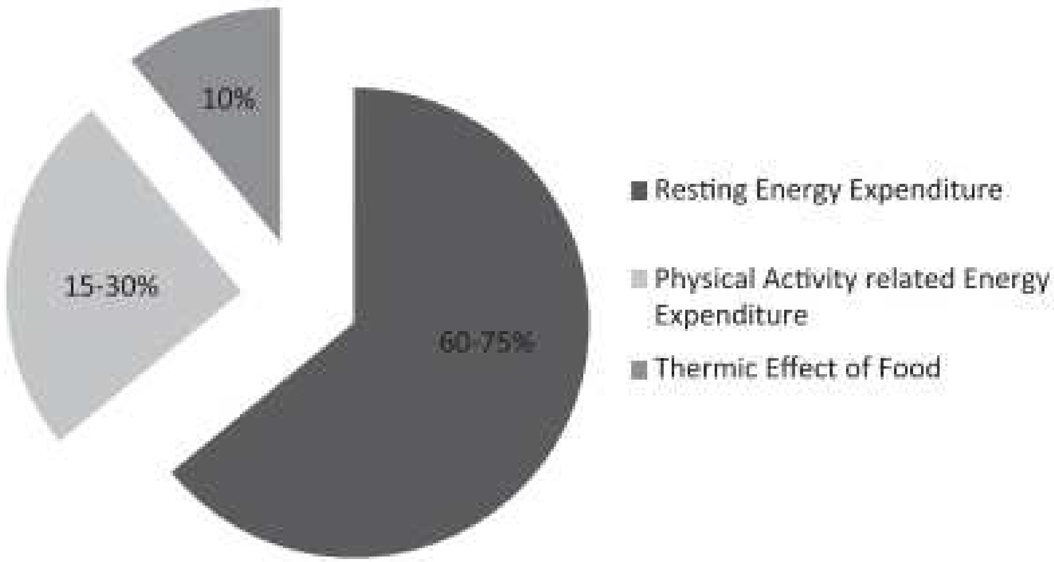
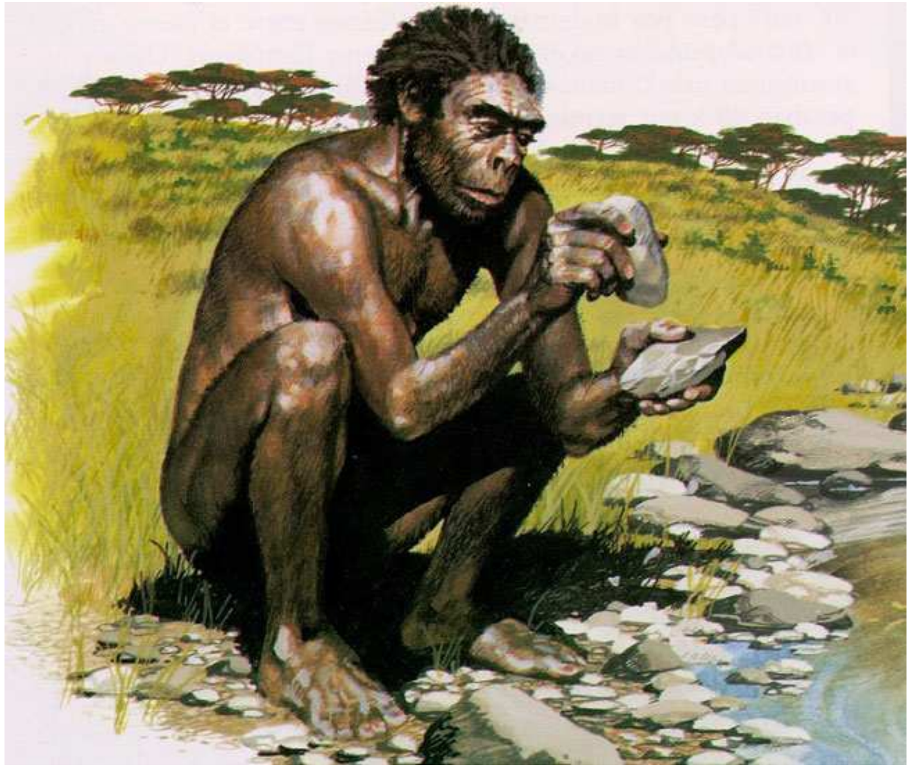
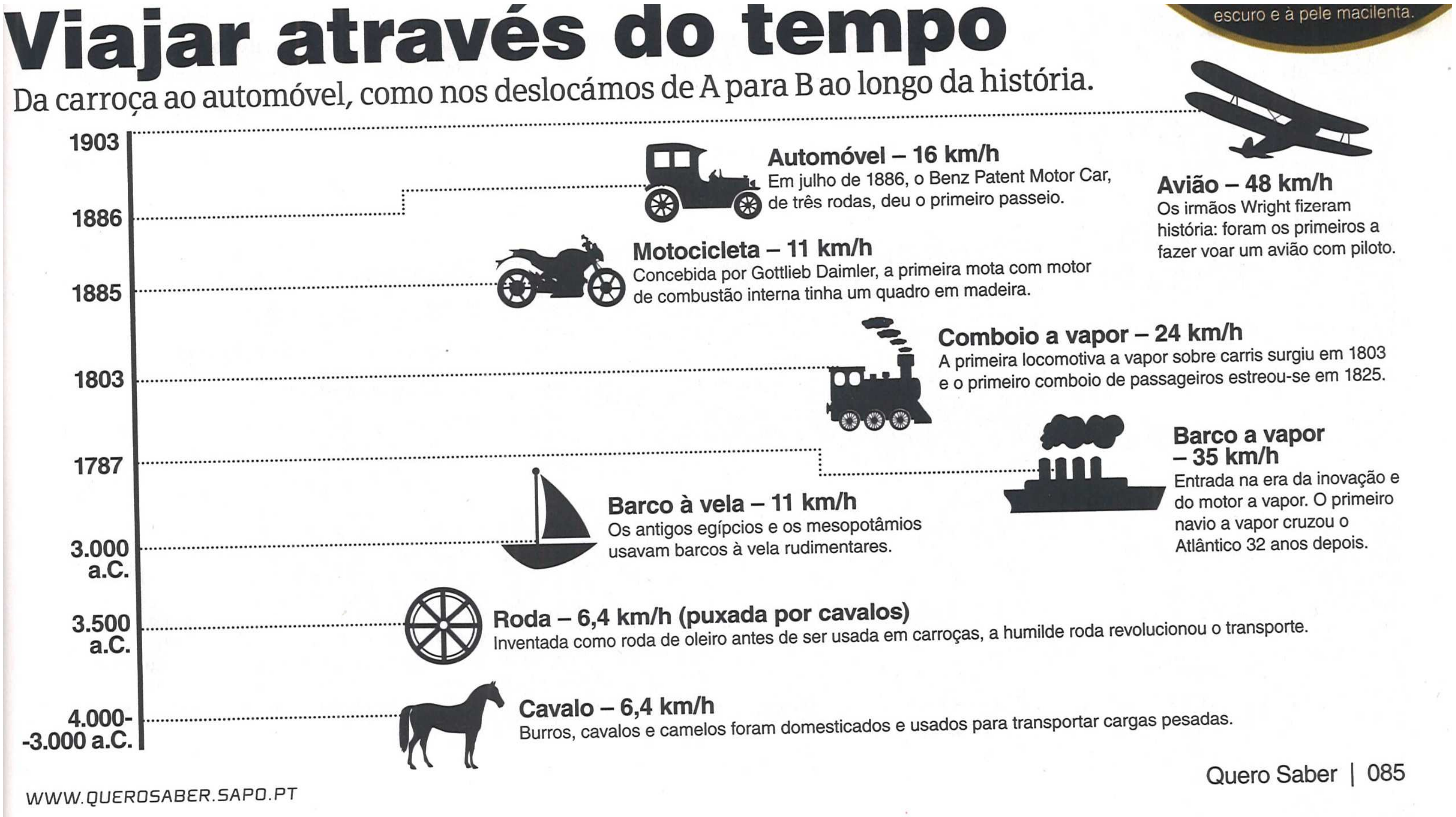


Figure 1. Components of typical total daily energy expenditure. Resting energy expenditure indicates the energy needed to maintain vital life functions during basal and sleeping conditions; physical activity–related energy expenditure, the energy needed to maintain movement demand above that of resting conditions; and thermic effect of food, the energy required for purposes of digestion and the breakdown of food stuff. Modified from McArdle et al.[21]

Combustion history



Combustion in transports



escuro e à pele macilenta.

Viajar através do tempo

Da carroça ao automóvel, como nos deslocámos de A para B ao longo da história.

1903 — **Avião – 48 km/h**
Os irmãos Wright fizeram história: foram os primeiros a fazer voar um avião com piloto.

1886 — **Automóvel – 16 km/h**
Em julho de 1886, o Benz Patent Motor Car, de três rodas, deu o primeiro passeio.

1885 — **Motocicleta – 11 km/h**
Concebida por Gottlieb Daimler, a primeira mota com motor de combustão interna tinha um quadro em madeira.

1803 — **Comboio a vapor – 24 km/h**
A primeira locomotiva a vapor sobre carris surgiu em 1803 e o primeiro comboio de passageiros estreou-se em 1825.

1787 — **Barco a vapor – 35 km/h**
Entrada na era da inovação e do motor a vapor. O primeiro navio a vapor cruzou o Atlântico 32 anos depois.

3.000 a.C. — **Barco à vela – 11 km/h**
Os antigos egípcios e os mesopotâmios usavam barcos à vela rudimentares.

3.500 a.C. — **Roda – 6,4 km/h (puxada por cavalos)**
Inventada como roda de oleiro antes de ser usada em carroças, a humilde roda revolucionou o transporte.

4.000--3.000 a.C. — **Cavalo – 6,4 km/h**
Burros, cavalos e camelos foram domesticados e usados para transportar cargas pesadas.

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Combustion in our lifes

- Cooking; Thermal comfort

Combustion of natural gas

Combustion of wood/pellets

Combustion in our lifes

- Transport

Combustion of hydrogen

Combustion of maritime diesel

Combustion of diesel fuel

Combustion of jet fuel

Combustion in our lifes

- Generation of electricity

Combustion of coal

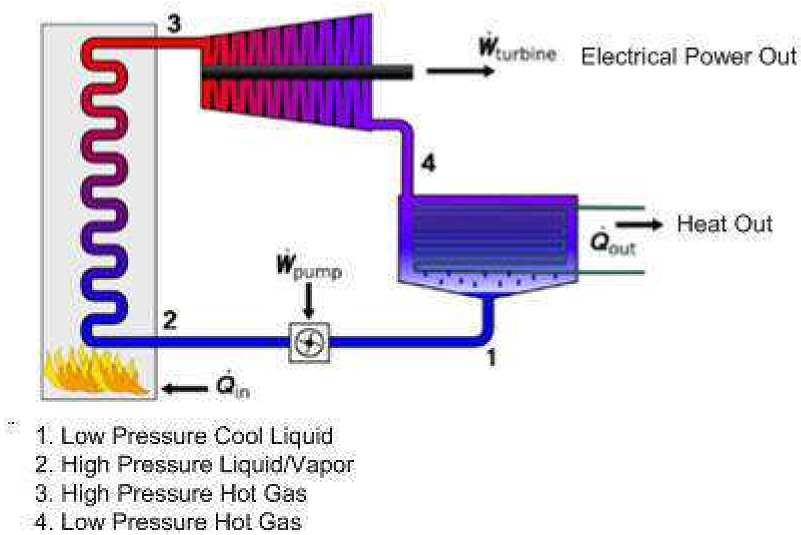
Combustion of diesel

Combustion of natural gas

Combustion in our lifes

- Generation of electricity

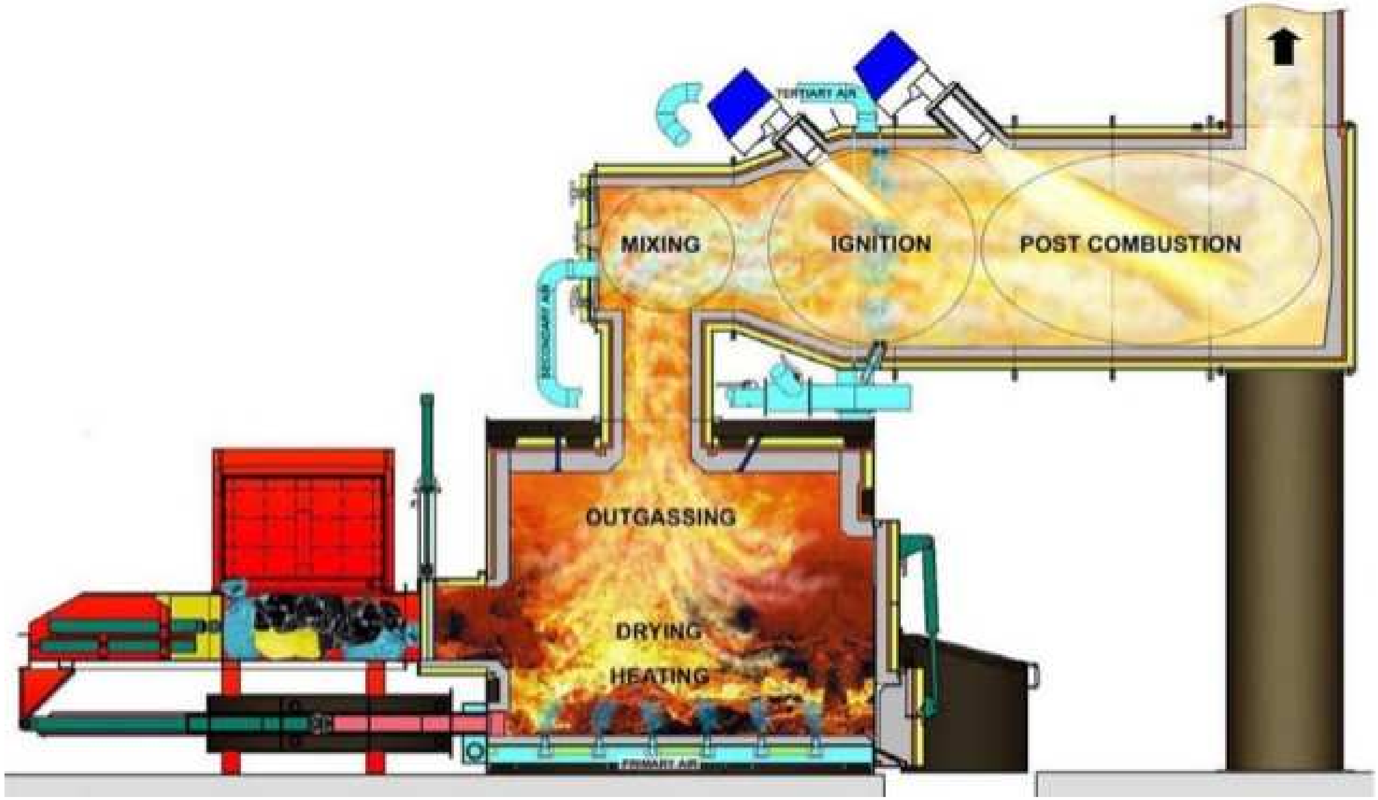
Combustion of coal



Combustion of natural gas

Combustion in our lifes

- Waste disposal



Combustion in our lifes

- Waste disposal up to 1000 kg/capita/year

Table 2.1 Key figures—municipal solid waste (kg/capita/year)

Area	Ref.	Waste generation [kg/cap./year] Range	Mean	Annual growth rate
OECD—total	/2/	263–864	513	1.9%
North America	/2/		826	2.0%
Japan	/2/		394	1.1%
OECD—Europe	/2/		336	1.5%
Europe (32 countries)	/3/	150–624	345	n.a.
8 Asian Capitals	/4/	185–1000	n.a.	n.a.
South and West Asia (cities)	/5/	185–290	n.a.	n.a.
Latin America and the Caribbean	/6/	110–365	n.a.	n.a.

Combustion in our lifes

- Waste disposal

Table 2.2 Composition of municipal wastes (percentage of wet weight)

% of waste		*Guangzhou, China, 8 districts*		*Manila*	*22 European Countries*	
	Year	*1993*		*1997*	*1990*	
	Ref.	*/7/*		*/9/*	*/3/*	
Fraction		*Range*	*Mean*	*Mean*	*Range*	*Mean*
Food and organic waste		40.1 – 71.2	46.9	45.0	7.2 – 51.9	32.4
Plastics		0.9 – 9.5	4.9	23.1	2 – 15	7.5
Textiles		0.9 – 3.0	2.1	3.5	n.a.	n.a.
Paper & cardboard		1.0 – 4.7	3.1	12.0	8.6 – 44	25.2
Leather & rubber		..	..	1.4	n.a.	n.a.
Wood		..	..	8.0	n.a.	n.a.
Metals		0.2 – 1.7	0.7	4.1	2 – 8	4.7
Glass		0.8 – 3.4	2.2	1.3	2.3 – 12	6.2
Inerts (slag, ash, soil, etc.)		14.0 – 59.2	40.2	0.8	..	..
Others		..	..	0.7	6.6 – 63.4	24.0

Notes: n.a. = Not applicable
.. = Negligible

@http://web.mit.edu/urbanupgrading/urbanenvironment/resources/references/pdfs/MunicipalSWIncin.pdf

Combustion in our lifes

- Waste incineration and energy generation

Osaka, Japan

Outputs	Efficiency	Use
Heat Only	Up to 80-90%[15] thermal efficiency.	Local district heating for buildings (residential, commercial) and or for industrial processes.
Electricity	14%-27%*	Can be supplied to national grid for sale and distribution.
Heat and Power	Dependent on specific demand for heat and power.	Combination of above.
* The lower efficiency performance is more typical of older facilities and it is possible that in the future the efficiency of electricity generation using incineration will increase.		

Table 3: Examples of Energy Efficiency for Incineration

https://www.gov.uk/government/uploads/system/uploads/attachment_data/file/221036/pb13889-incineration-municipal-waste.pdf

Combustion in our lifes

- Waste incineration and energy generation

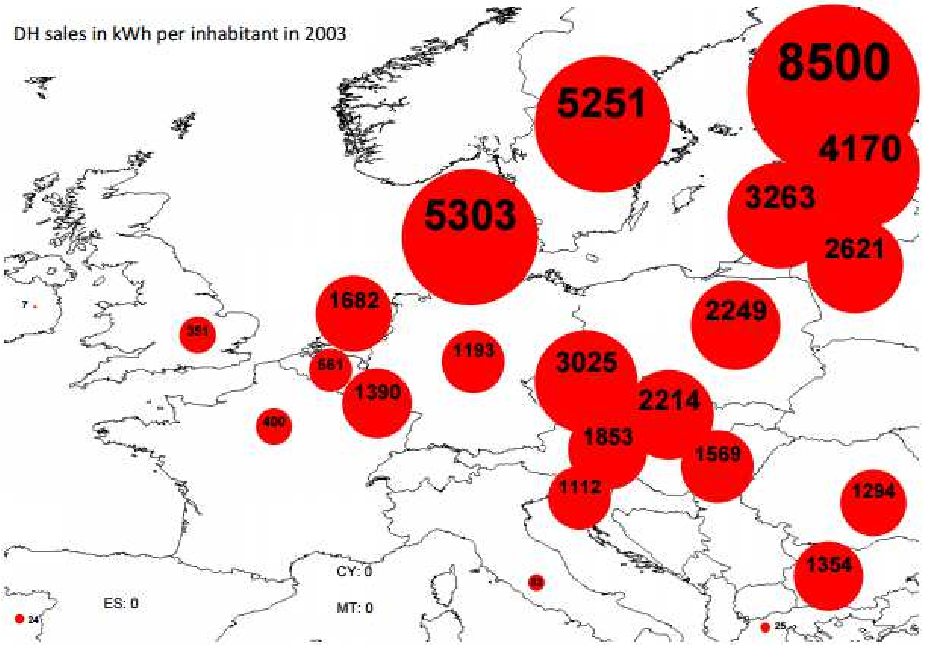
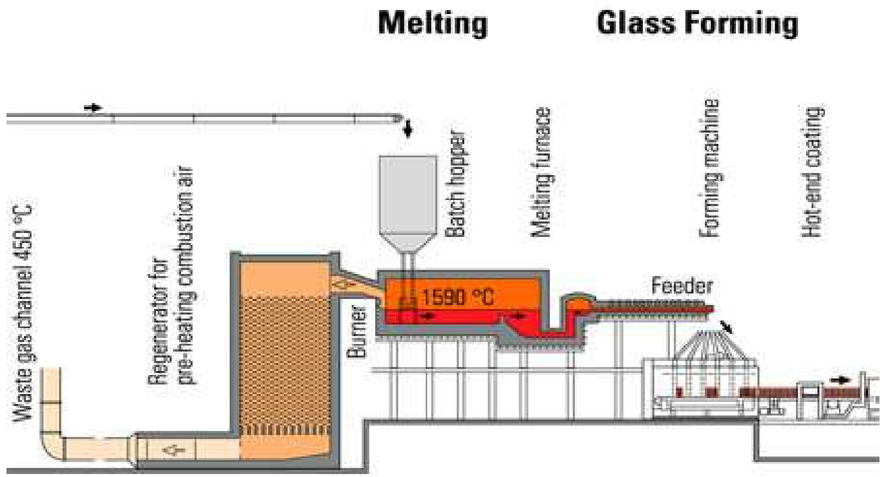


Figure 3.1: Heat sales per inhabitant from DH (District Heating) networks in the EU 27 Member States in 2003 (Data from EcoHeatCool report). The surface area of the circles is proportional to the figures in kWh

@http://iet.jrc.ec.europa.eu/remea/sites/remea/files/r1_climate_factor_report_final.pdf

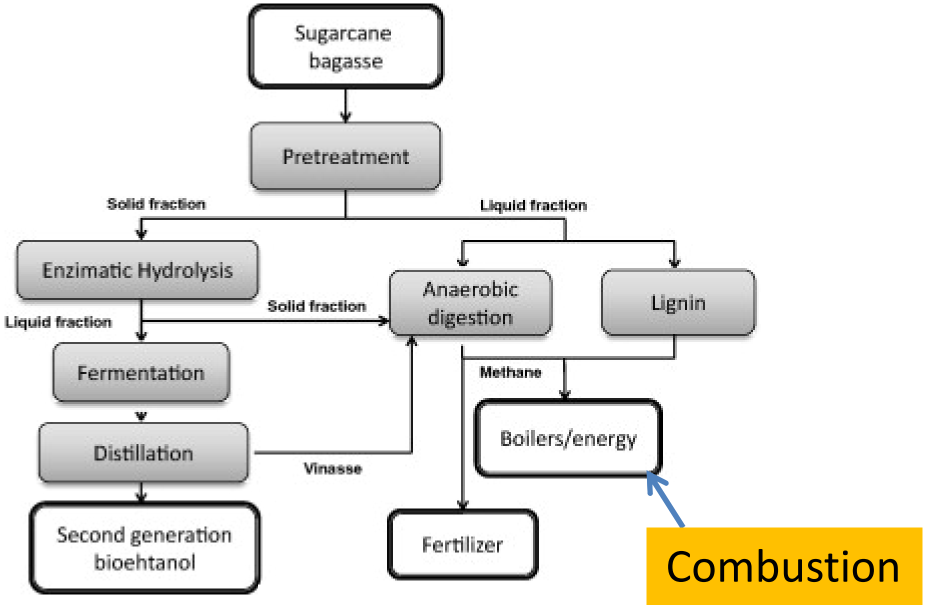
Combustion in our lifes

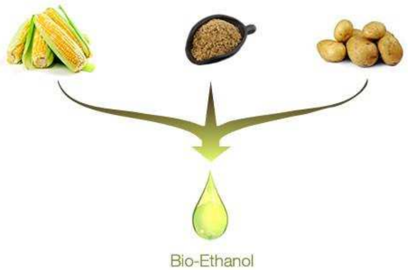
- Product manufacturing

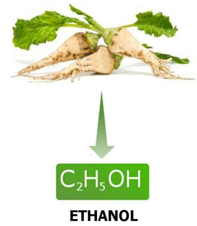


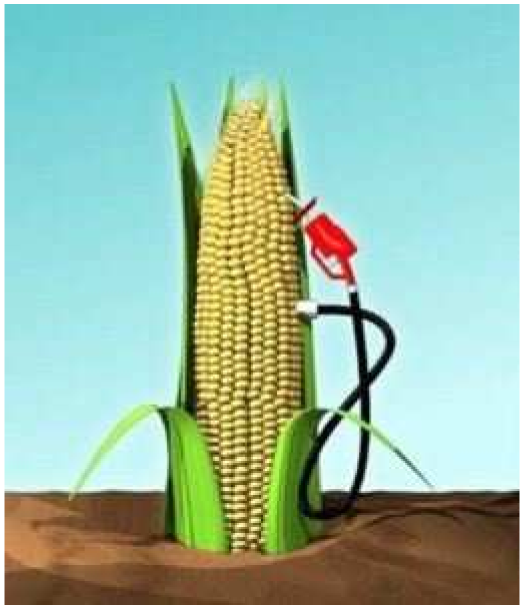
Combustion in our lifes

- Product manufacturing









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Combustion in our lifes

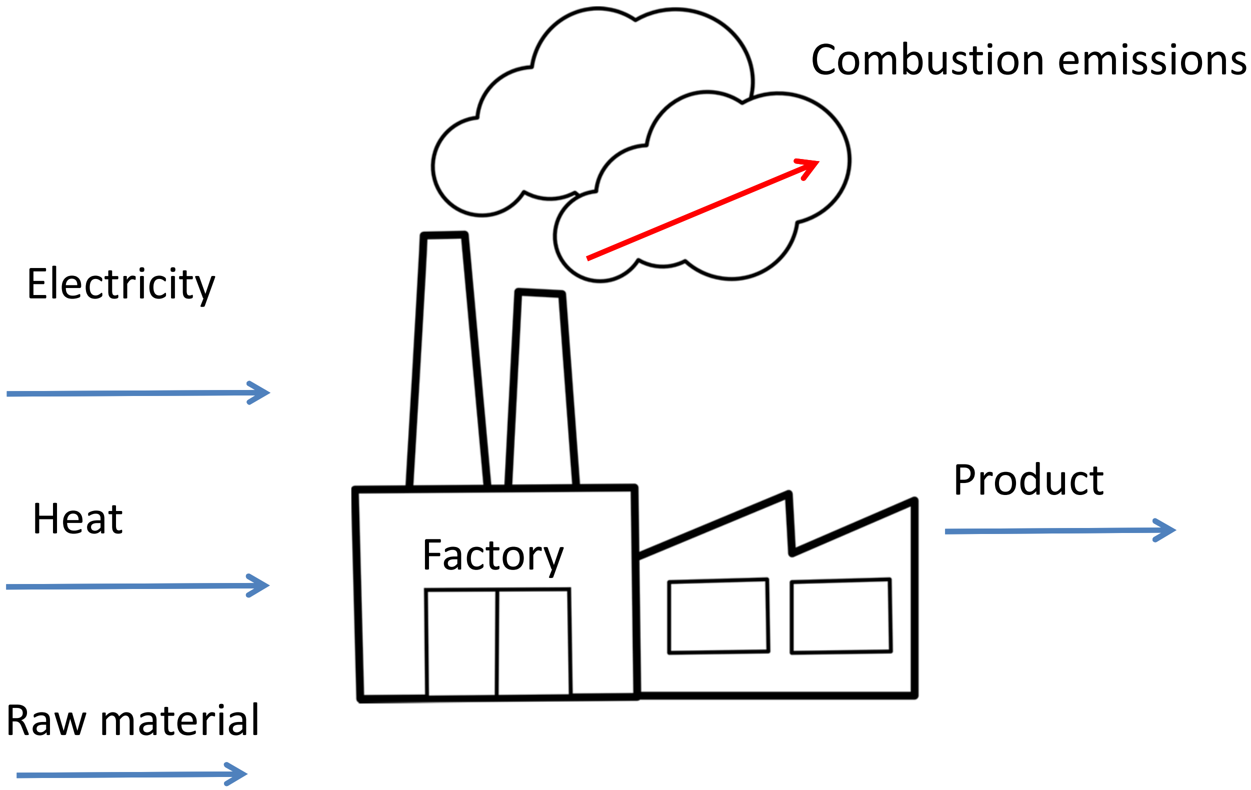
- Product manufacturing

SESEC
Sustainable Energy Saving
for the European Clothing Industry

Energy can either be supplied by direct combustion of gas in the productive machines, or indirectly through production of steam or hot water

Combustion in our lifes

- Product manufacturing



Combustion in our lifes

- Product manufacturing

Combustion

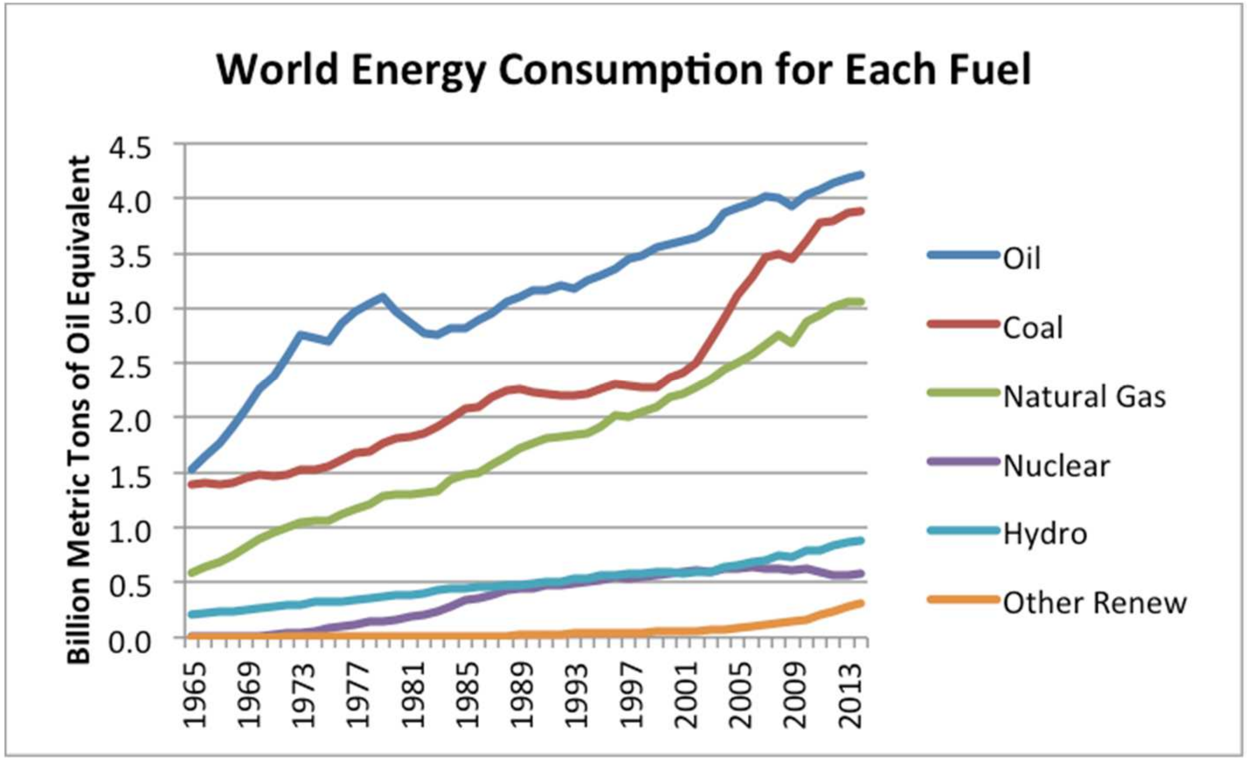
robots welding in an automobile factory (automotive, manufacturing, automation)

Combustion in our lifes

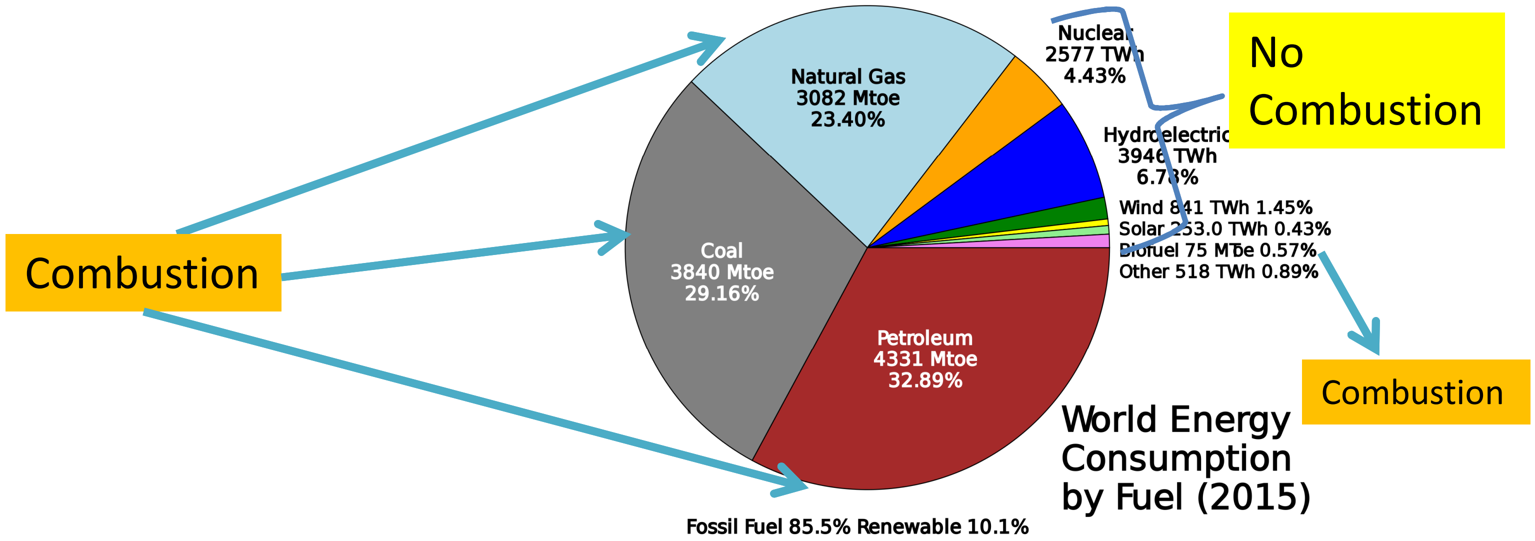
Combustion in our lifes

Combustion in our lifes

Combustion use in the world

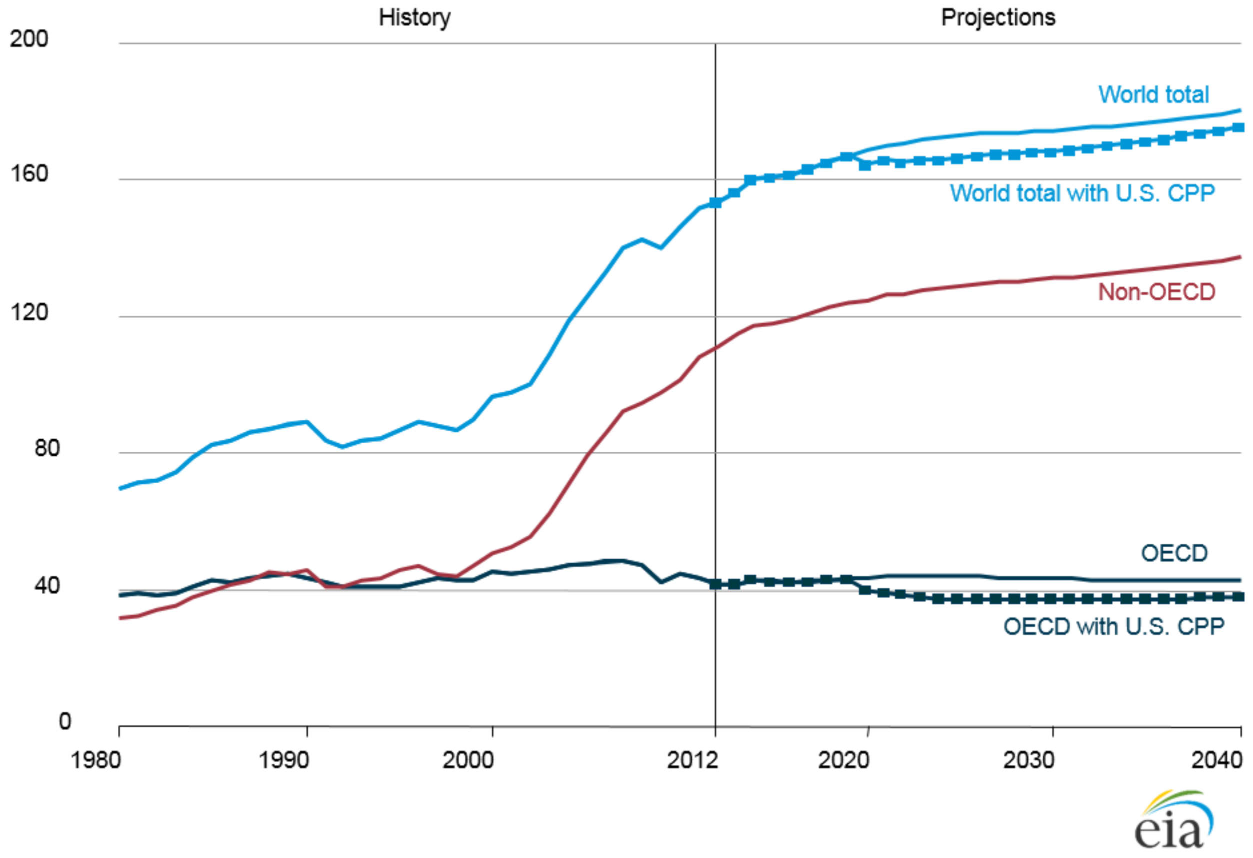


Combustion use in the world



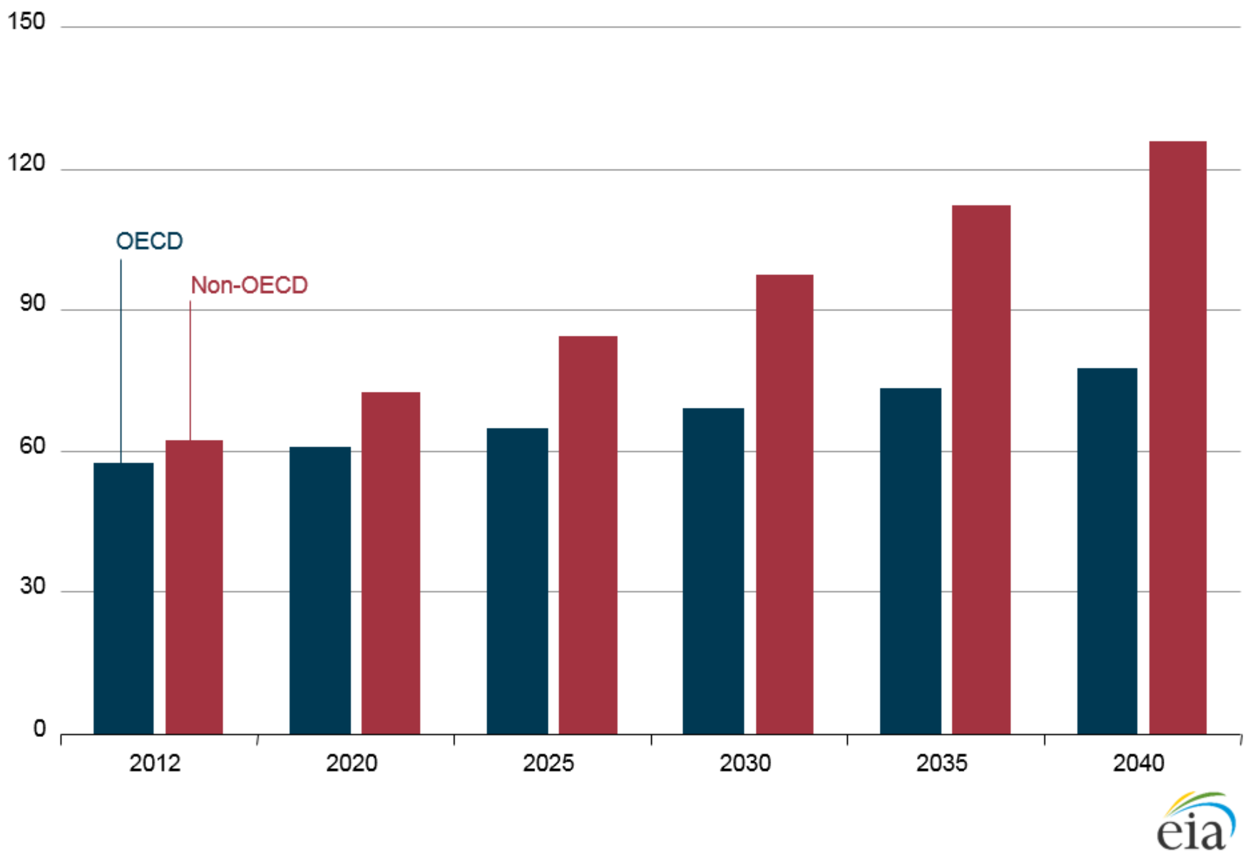
Combustion use in the world

Figure 4-1. World coal consumption by region, 1980–2040
quadrillion Btu

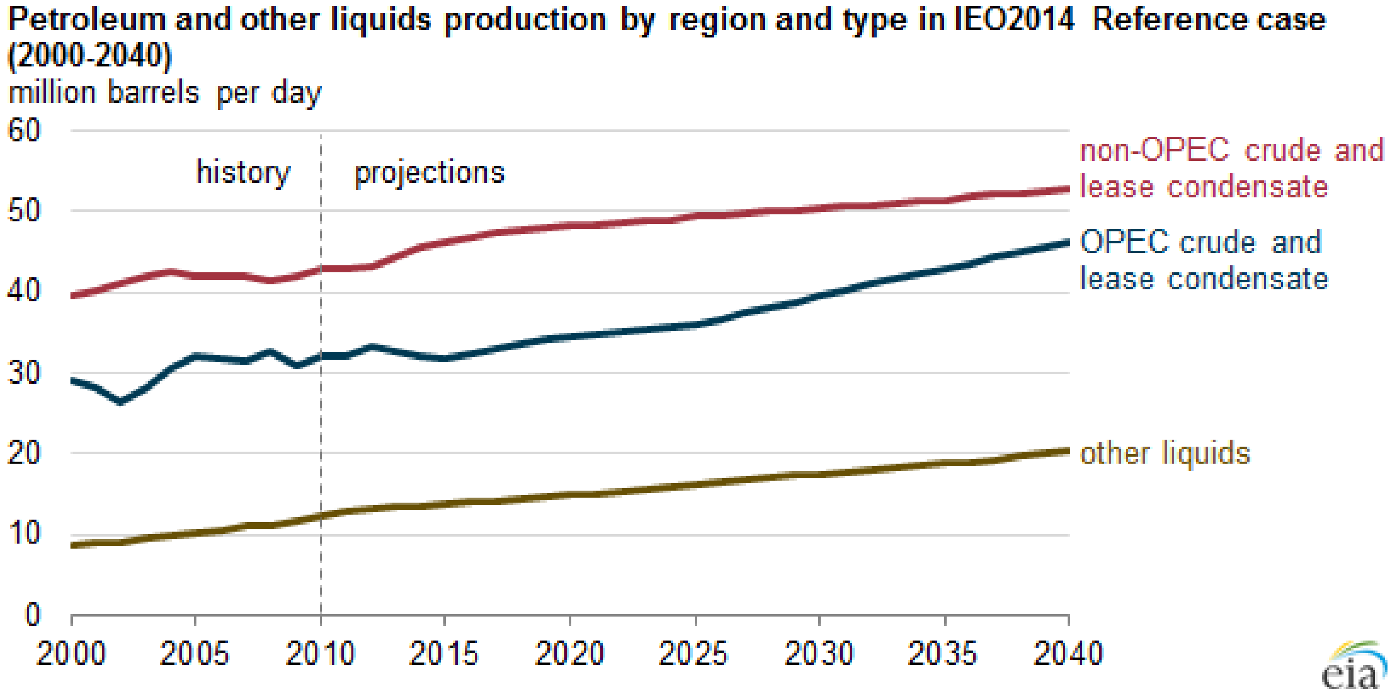


Combustion in the world

Figure 3-1. World natural gas consumption, 2012–40
trillion cubic feet



Combustion in the world



Combustion problem

1 toe = 41.868 GJ or 11.63 MWh

the amount of energy released by burning one tonne of crude oil

Combustion problem

Why do we have emissions???

Products of ideal combustion

$$C_nH_m + S + N_2 + O_2 \rightarrow CO_2 + H_2O + N_2 + O_2$$
$$+ NO_x + CO + SO_x + Soot + UHC$$

Fuel

Air

Products of non-ideal combustion

NO_x: Affects ozone (O_3) concentration
CO_2: Absorbs outgoing infrared radiation
CO: Toxic
Soot: Visible

Combustion problem

1 toe = 41.868 GJ or 11.63 MWh

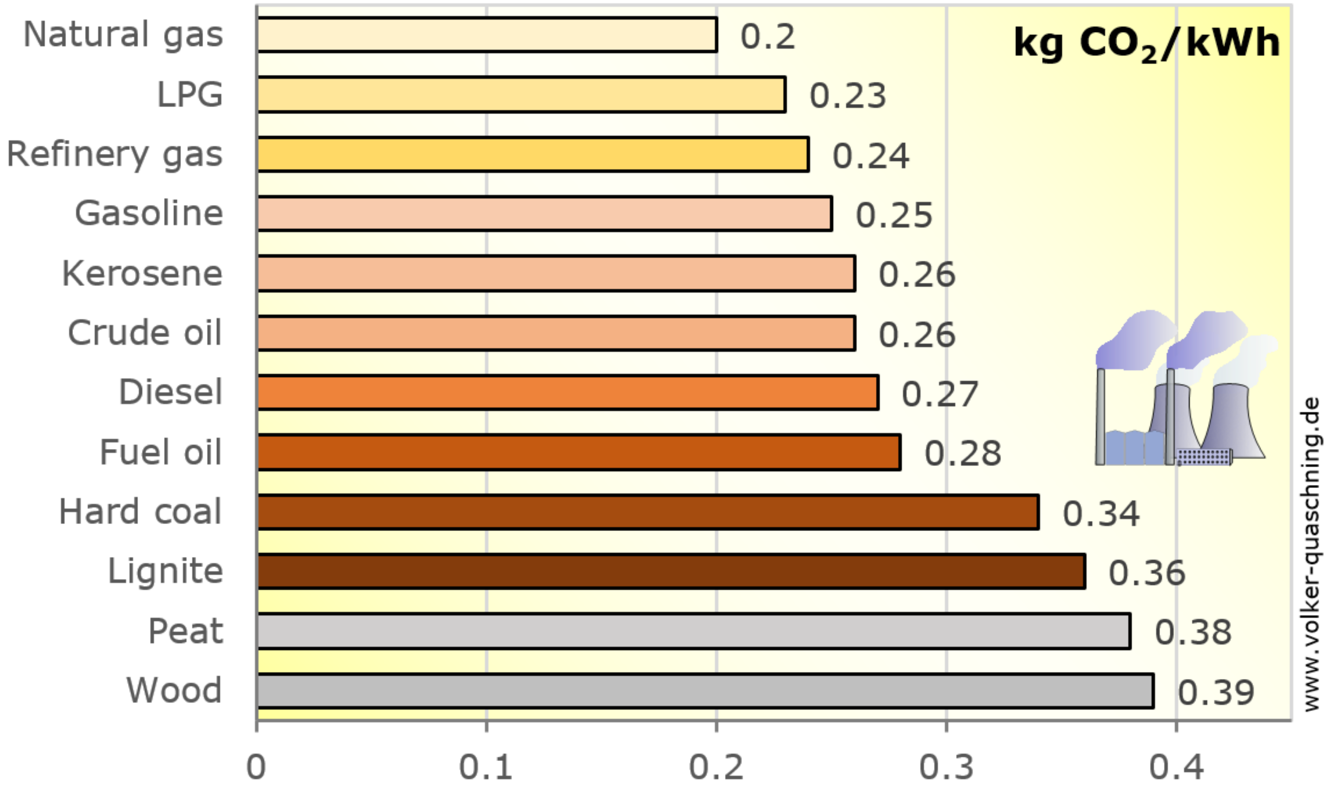
the amount of energy released by burning one tonne of crude oil

1 toe ~ 2.5 tCO_2

Table 1.3 Typical pollutants of concern from selected sources

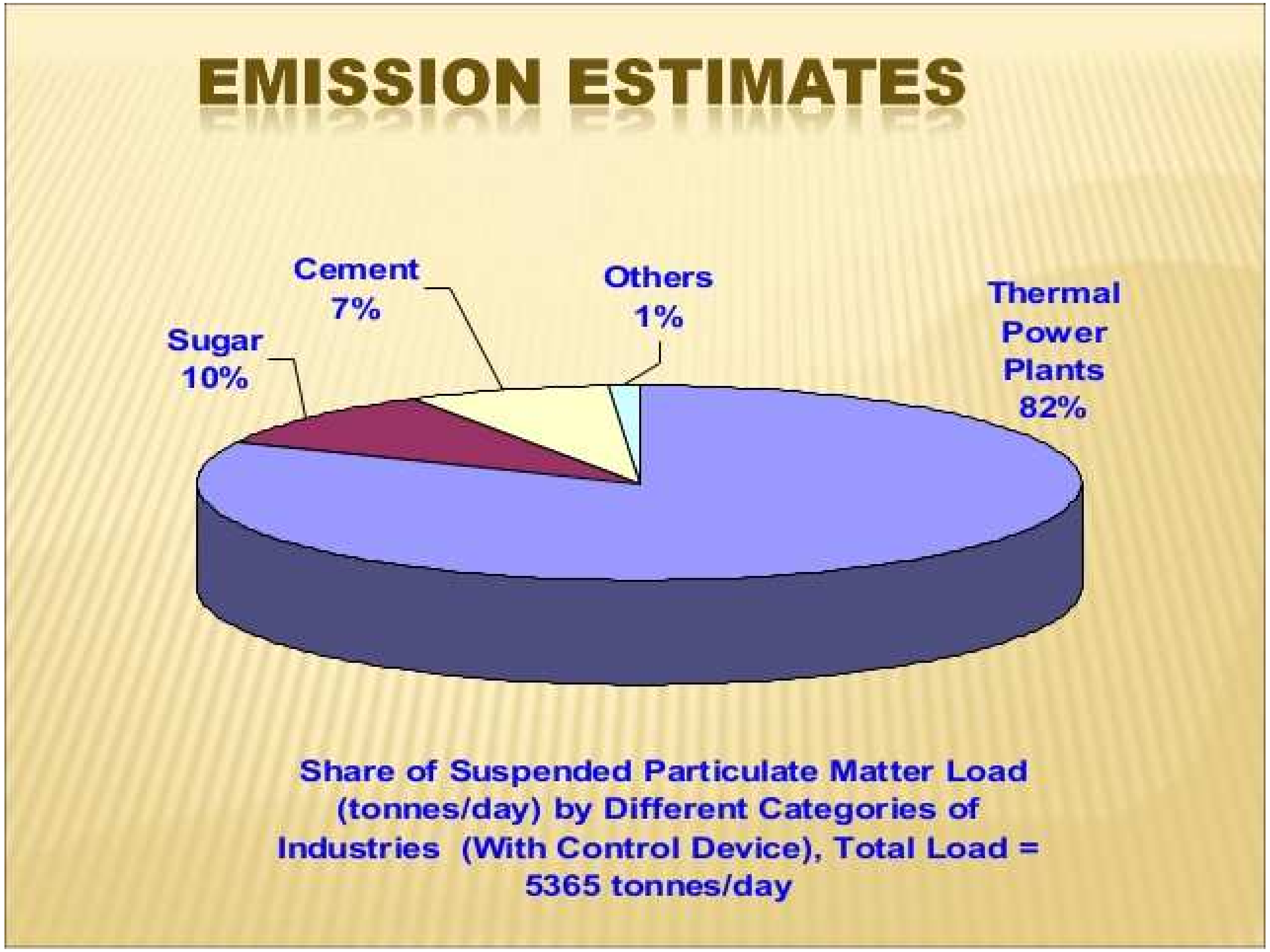
Source	Pollutants				
	Unburned Hydrocarbons	Oxides of Nitrogen	Carbon Monoxide	Sulfur Oxides	Particulate Matter
Spark-ignition engines	+	+	+	−	−
Diesel engines	+	+	+	−	+
Gas-turbine engines	+	+	+	−	+
Coal-burning utility boilers	−	+	−	+	+
Gas-burning appliances	−	+	+	−	−

Combustion problem



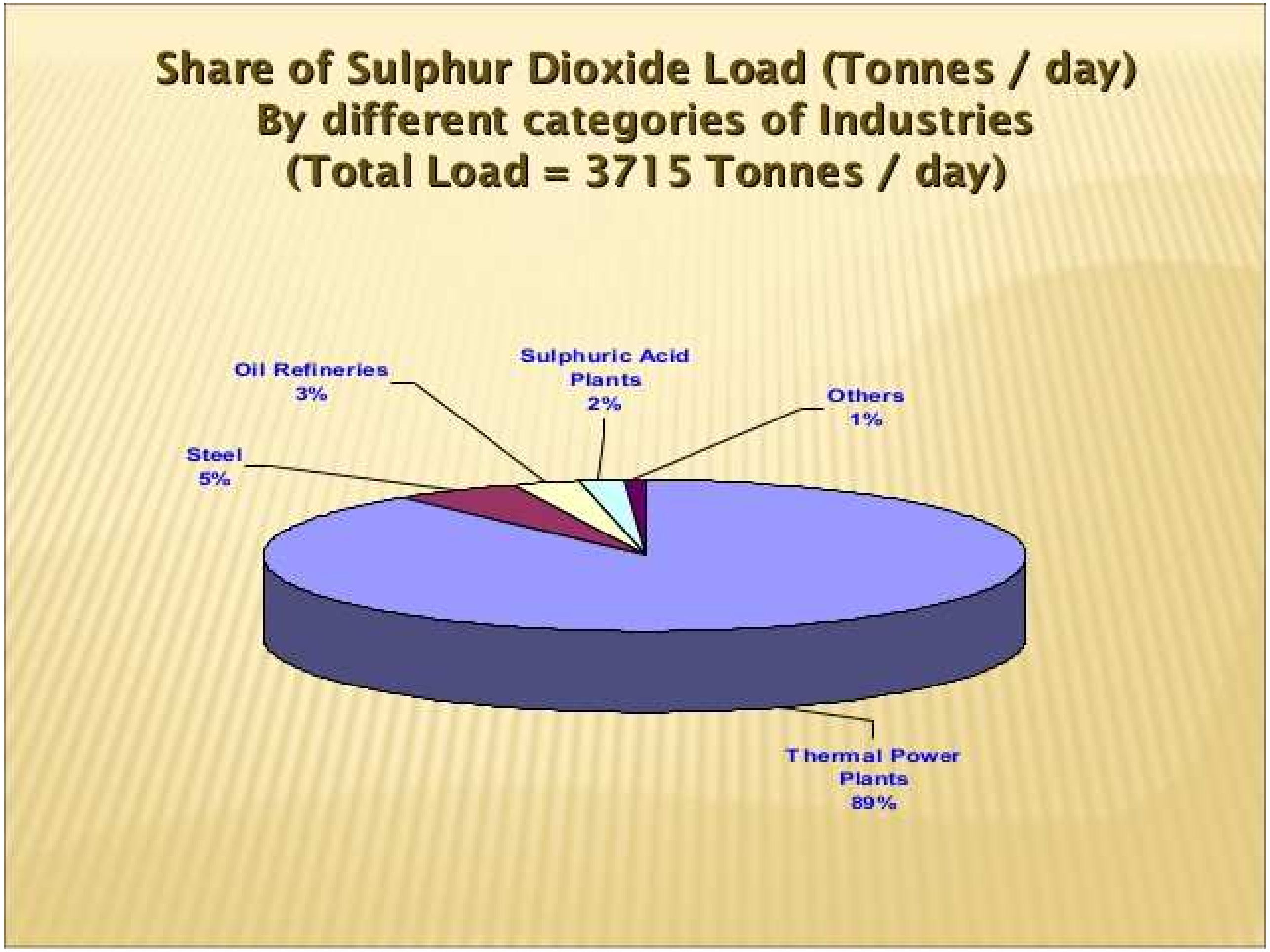
Specific Carbon Dioxide Emissions of Various Fuels

Combustion problem



PM10;PM2.5

Combustion problem



SO_2

Combustion problem

Transports, local emissions

LOCAL AIR QUALITY EMISSIONS (EUROPE)

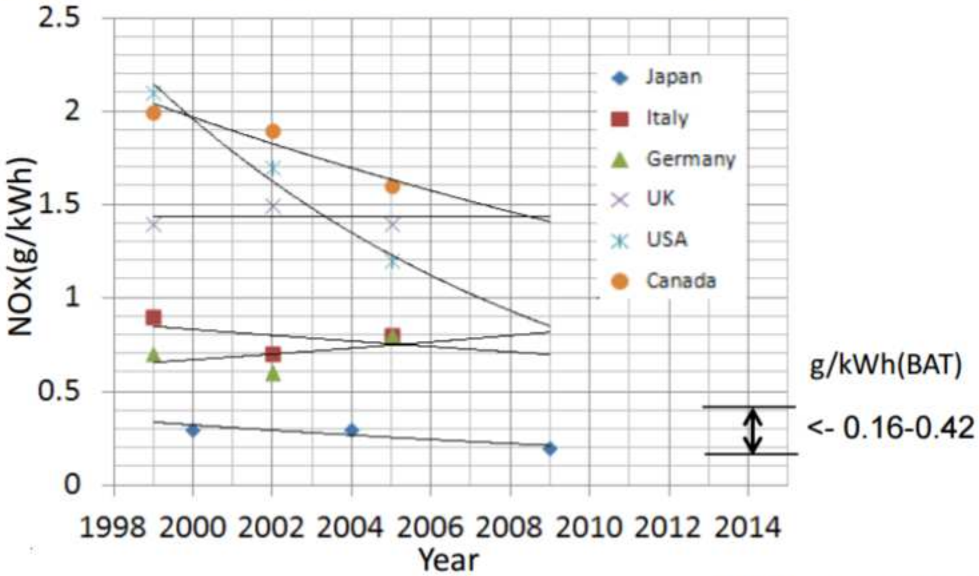
					NON-TRANSPORT
NOx	32.9%	0.9%	19.1%	4.5%	42.6%
CO	26.6%	0.2%	2.3%	0.7%	70.2%
SOx	0.1%	0.0%	20.9%	0.5%	78.5%
VOLATILE ORGANIC COMPOUNDS	15.4%	0.14%	2.52%	0.40%	81.54%
FINE PARTICLES (PM2.5)	14.2%	0.4%	11.4%	0.6%	73.4%

In % of total emissions | source: European Environment Agency, 2013

THANKS TO IMPROVEMENTS IN AIRCRAFT TECHNOLOGY, THE IMPACT OF AIRCRAFT EMISSIONS ON LOCAL AIR QUALITY IS RELATIVELY LOW COMPARED TO OTHER SOURCES.

Combustion problem

Figure: Trend of average NOx emission from thermal power plants



BAT :Best Available Technology

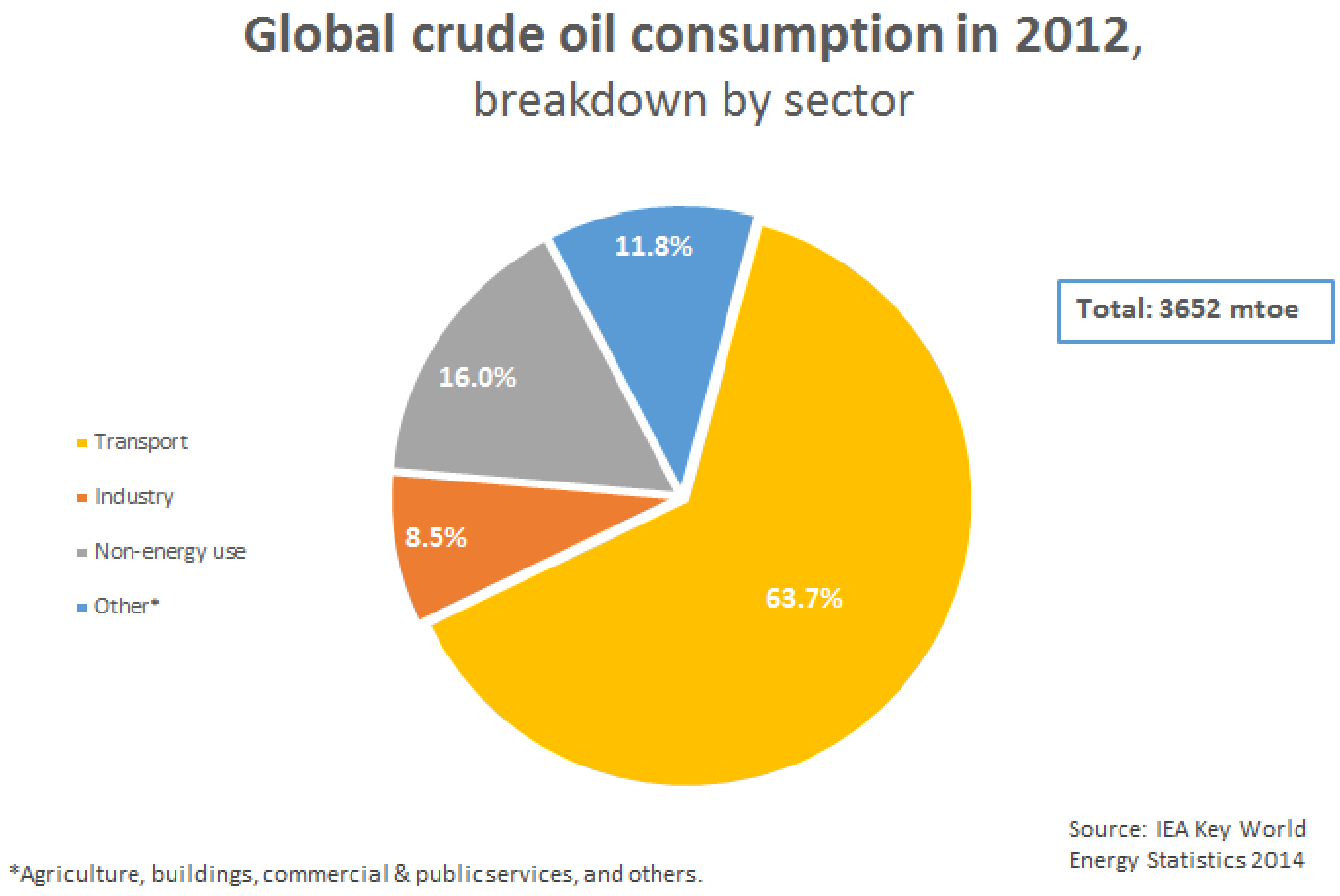
Source: Federation of Electric Power Companies,INFOBASE2010

NO_2, NO

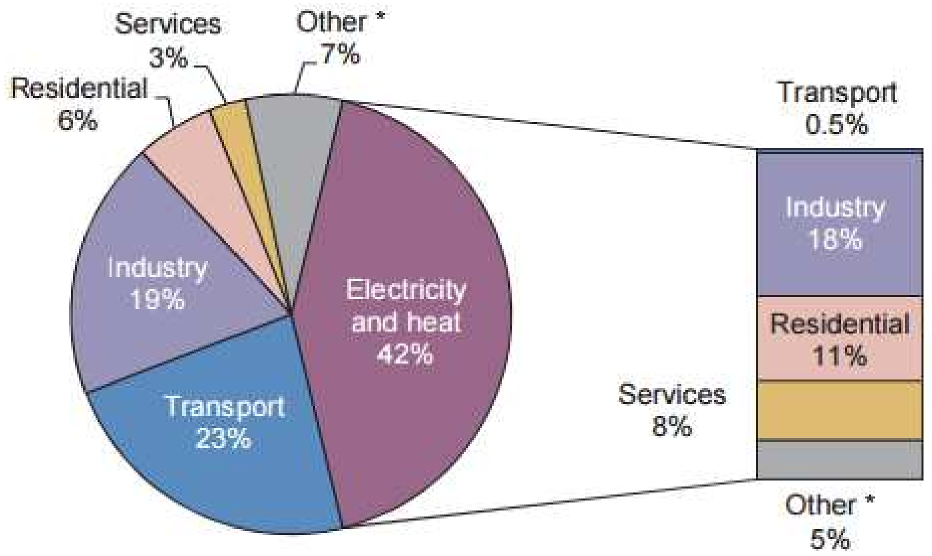
Typically called

NOx

Combustion in the world



Combustion in the world



Note: Also shows allocation of electricity and heat to end-use sectors.

* Other includes agriculture/forestry, fishing, energy industries other than electricity and heat generation, and other emissions not specified elsewhere.

World CO_2 emissions by sector in 2013 Note: Also shows allocation of electricity and heat to end-use sectors. * Other includes agriculture/forestry, fishing, energy industries other than electricity and heat generation, and other emissions not specified elsewhere.

Future goals

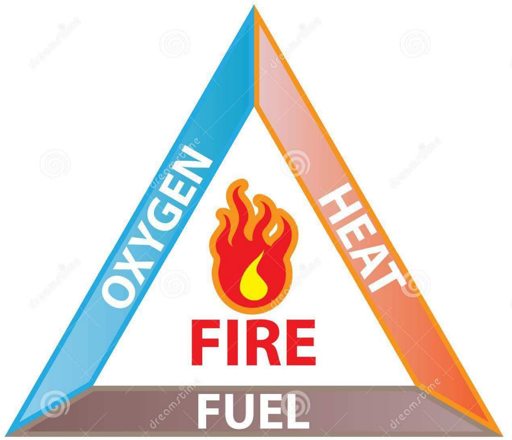
- Increase efficiency
- Reduce emissions

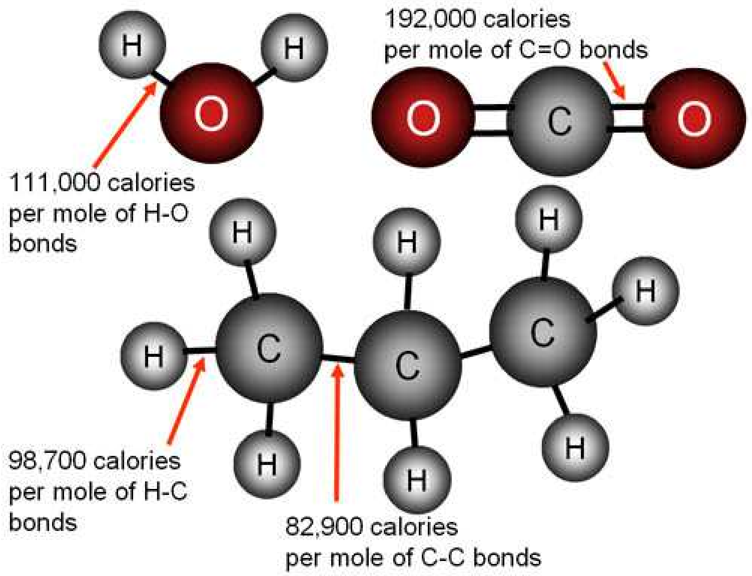
Avoid combustion in some sectors, e.g., transport

Combustion definition

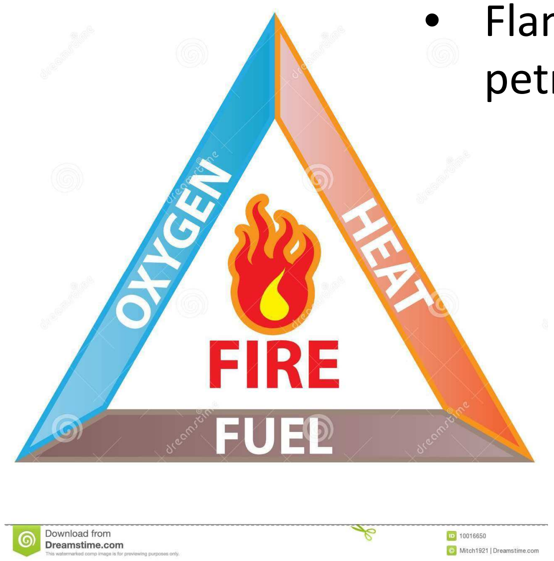
Combustion is essentially burning, fuels react with oxygen to release energy (heat or both light and heat)

Combustion transforms energy stored in chemical bonds to heat that can be utilized in a variety of ways

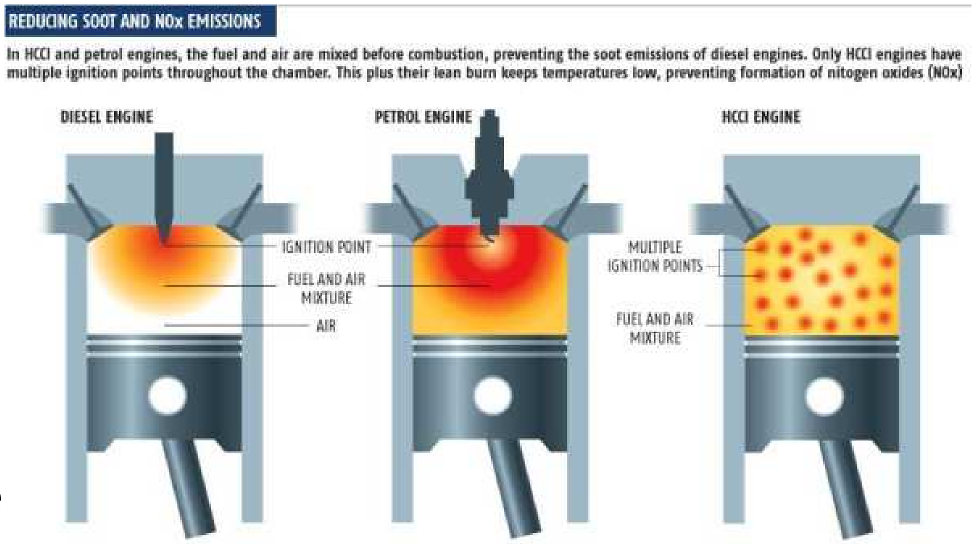




Combustion modes



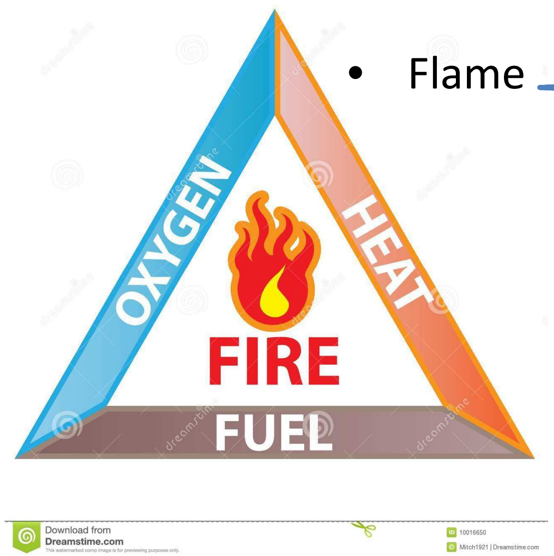
- Flame e.g. petrol engine



- Nonflame

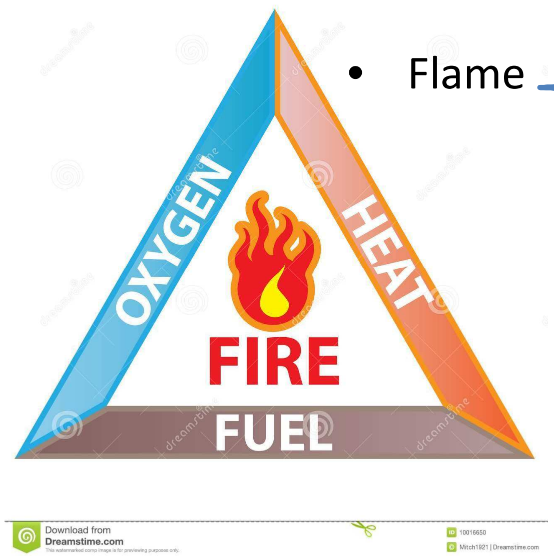
e.g. Diesel engine **autoignition** or homogeneous charge compression by high pressure

Combustion modes



- Flame
  - Premixed
  - Nonpremixed (difusion)
- Nonflame

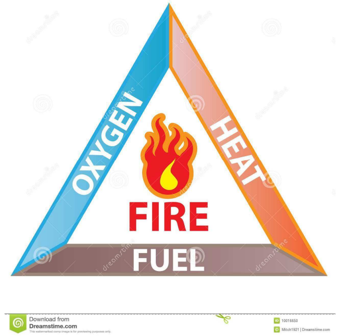
Flame types



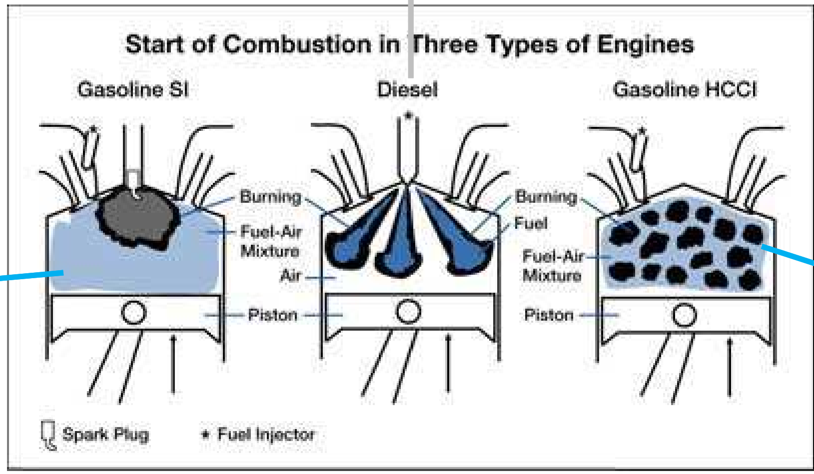
- Flame
  - Premixed

    Fuel and oxidizer are mixed at the molecular level prior to the occurrence of any significant chemical reaction
  - Nonpremixed (difusion)
    - Laminar
    - Turbulent
- Nonflame

Flame types



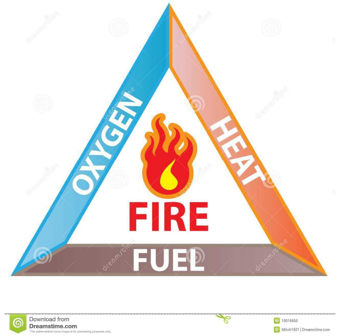
Nonflame combustion
Nonpremixed (difusion) flame

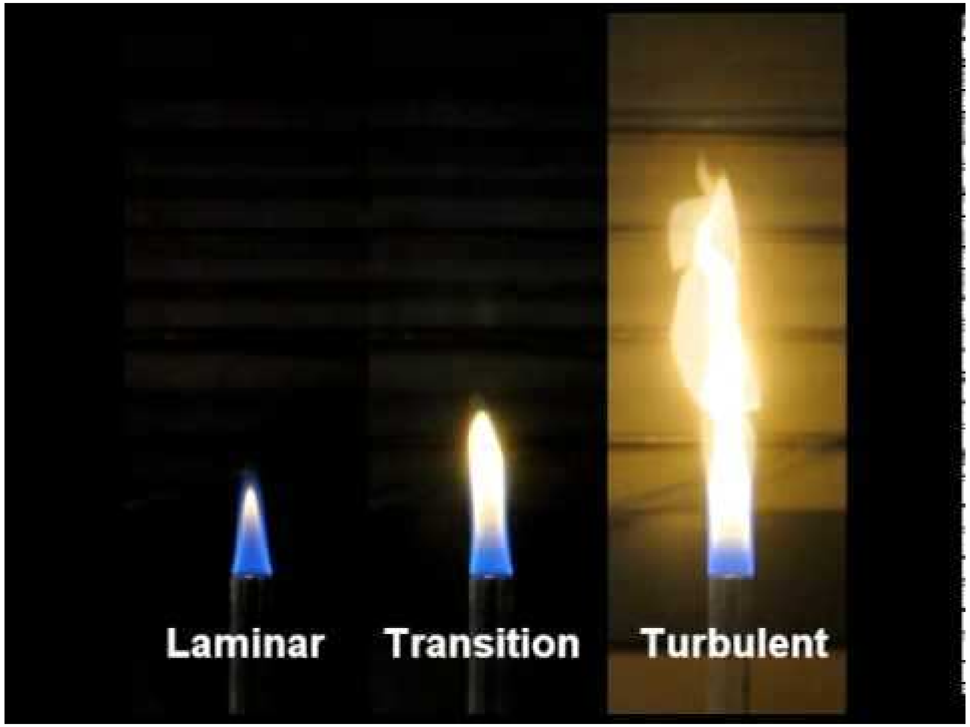


Flame combustion
Premixed flame

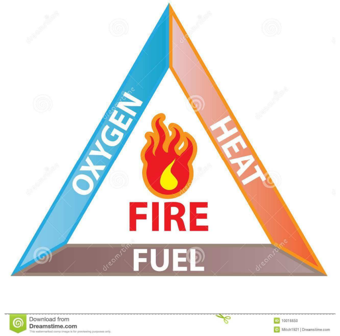
Nonflame combustion
Premixed flame

Flame types

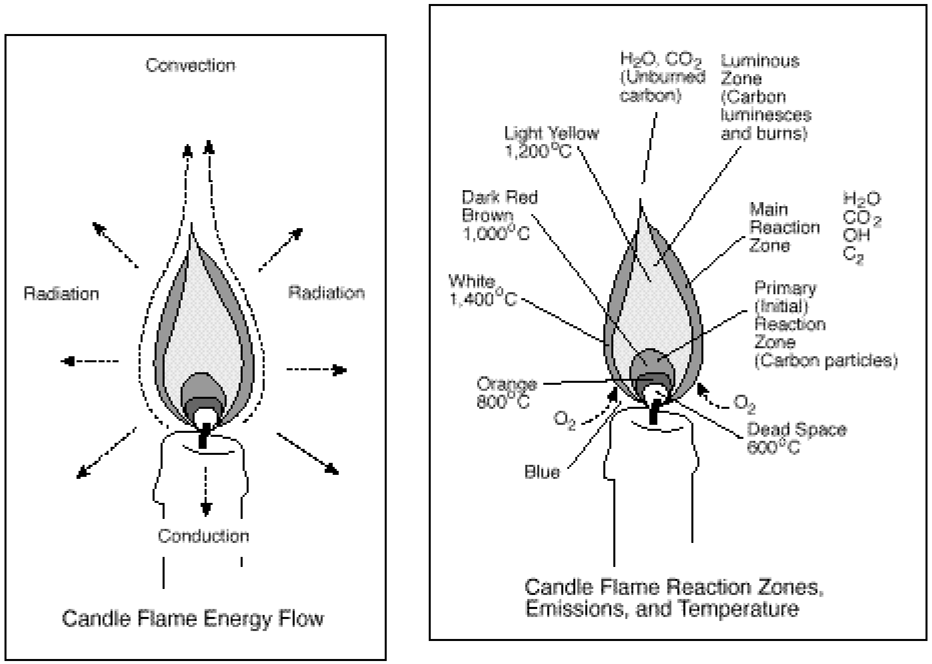




Flame types

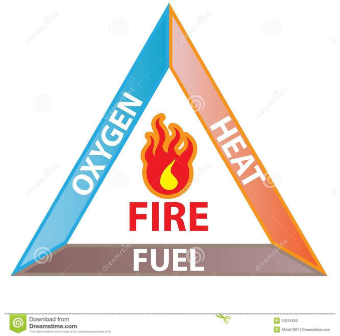


Candle
Fuel: wax
Oxidizer: air
Reaction zone between wax vapours and air (difusion flame)



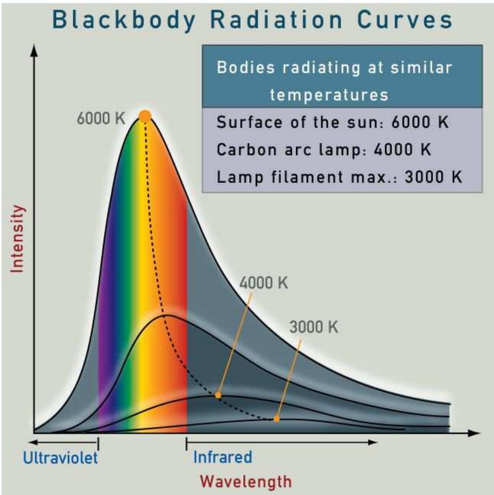
Candle flame diagrams adapted from "The Science of Flames" poster, National Energy Foundation, Salt Lake City, UT.

Flame types



Gravity

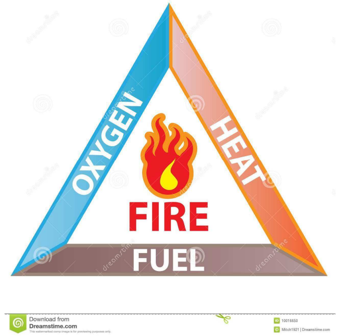
No gravity



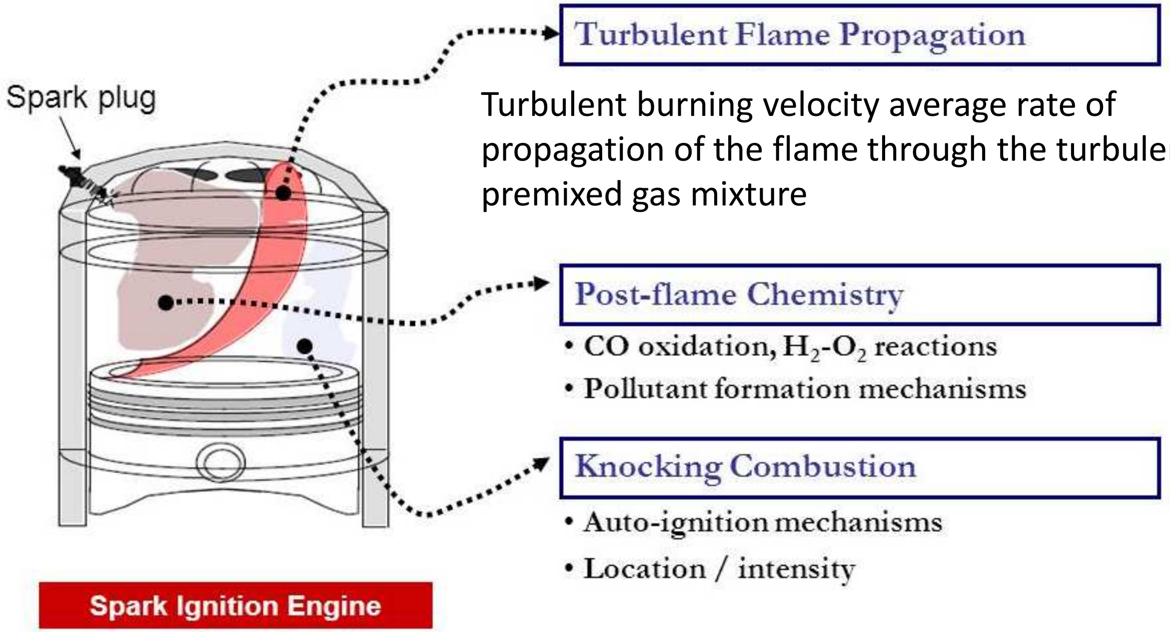
1'000 K	Red
1'500 K	Reddish orange
2'000 K	Yellowish orange
2'800 K	Yellow
3'500 K	Yellowish white
4'500 K	Warm white
5'500 K	White

Increasing temperature

Flame types



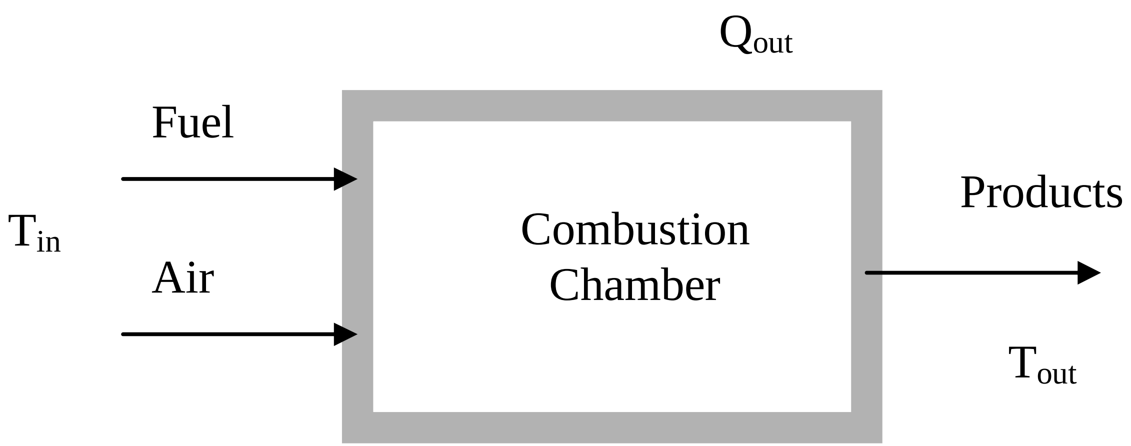
Spark ignition engine
Fuel: gasoline, ethanol, GPL or natural gas
Oxidizer: air
Reaction zone between premixture and air (premixed flame)



Turbulent burning velocity average rate of propagation of the flame through the turbulent premixed gas mixture

Thermochemistry

$$fuel + oxidizer \rightarrow products$$



Thermochemistry

$$fuel + oxidizer \rightarrow products$$

Generic hydrocarbon

$C_xH_yO_z$

O_2 or air

CO_2, H_2O and N_2

Air composition

21% O2 and 79% N2 (by volume)

Thermochemistry

$$fuel + oxidizer \rightarrow products$$

Air composition

21% O2 and 79% N2 (by volume)

$$C_xH_yO_z + n(0.21O_2 + 0.79N_2) \rightarrow xCO_2 + y/2H_2O + 0.79 * nN_2$$

$\lambda = \frac{\frac{A}{F}}{\left(\frac{A}{F}\right)_S}$ excess air coefficient

λ<1 No sufficient air; fuel is not completely burned

λ=1 Exact amount air, fuel is completely burned

λ>1 Excess air; fuel is completely burned

$\phi = \frac{1}{\lambda}$ equivalence ratio

$e(\%) = \frac{\frac{A}{F} - \left(\frac{A}{F}\right)_S}{\left(\frac{A}{F}\right)_S} * 100\% = \frac{1-\phi}{\phi}$ = excess air in %

Thermochemistry

$$\underbrace{fuel + oxidizer}_{\text{Reactants}} \rightarrow products$$

Air composition

21% O2 and 79% N2 (by volume)

$$C_xH_yO_z + n(O_2 + 3.76N_2) \rightarrow xCO_2 + y/2H_2O + 0.79 * nN_2$$

Right amount of oxidizer to burn all fuel?

1) Determine n,

2n+z=2x+y/2 ⇔ n=x+y/4-z/2

2) Determine mass air/mass fuel $(A/F)_s$ this is the stoichiometric air fuel ratio

$$\frac{mass\ air}{mass\ fuel} = \frac{n * (M_{O2} + 3.76M_{N2})}{Mfu} = \frac{(x + \frac{y}{4} - z/2) * (M_{O2} + 3.76M_{N2})}{Mfu}$$

Thermochemistry

P#1 Consider the combustion of CH_4 in air, determine:

a) The stoichiometric ratio;

b) Molar fractions of combustion products with 15% excess air.

Thermochemistry

P#2 A combustion chamber burns propane, $C3H8$ with excess air. Dry analysis (excuding water) of combustion products was: 2%$O2$, 12.4% $CO2$ and 85.6% $N2$. Determine:

a) The excess air.
b) The coefficient of air excess.
c) The equivalence ratio.

Thanks

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